

Lab #6: Molecular Shapes**Objectives:**

1. To become familiar with the three-dimensional shapes of molecules and polyatomic ions.
2. To draw electron-dot structures for simple molecules and polyatomic ions.
3. To build molecular models from electron-dot structures drawn and applying VSEPR model.
4. To predict the polarity and the type of intermolecular forces involved in molecules from its shape.

Pre-lab:

Copy the observation table. Draw the electron dot structures for the molecules or ions. Show all your work.

Materials:

Molecular Model Kit

Procedure:

1. From the electron dot structures drawn in the pre-lab section and applying VSEPR model, build the three-dimensional structure of the molecules or polyatomic ions using either the Molecular Model Kit.
2. Draw (using --- , --- , --- , and indicate all lone pairs) the three-dimensional structures built.
3. Name the three-dimensional structures built.
4. Determine if the molecule is polar or non-polar for each of the molecule along with the types of intermolecular forces involved.

Observations:

	Compounds or Ions	Electron-Dot Structures	Sketch of 3-D Model	Name of Shape	Polar or Non-polar	Type(s) of Intermolecular Force
1.	HCl					
2.*	SiH ₄					
3.*	PH ₃					
4.	HCN					
5.	CH ₃ F					
6.*	H ₂ S					
7.	CH ₂ O					
8.*	O ₃					
9.	H ₂ O ₂					
10.	CO ₃ ²⁻					
11.	NO ₂ ¹⁻					
12.	NO ₂					
13.*	SO ₂					
14.*	PF ₅					
15.*	SCl ₆					
16.*	PO ₄ ³⁻					
17.*	SO ₄ ²⁻					
18.*	SCl ₄					
19.*	IF ₂ ¹⁻					
20.*	BrF ₅					
21.*	XeCl ₄					
22.*	BrCl ₃					

Analysis:

1. Explain the polarity of each molecule from the observation table.
2. Three of the compounds from the observation table have their boiling points listed below. Account for the differences in these boiling points.

Molecule	Boiling Point
CH ₃ F	-78°C (195 K)
H ₂ O ₂	150°C (423 K)
SiH ₄	-107°C (166 K)

Evaluation:

1. Why is it difficult to draw the Lewis Dot Diagram for NO₂? How does NO₂¹⁻ solve this problem? Besides NO₂¹⁻, describe other ways to resolve this difficulty?
2. Can a neutral NO₃ molecule exist? If not, what can be done so that one nitrogen atom and three oxygen atoms can form this molecule?
3. Most of the time, phosphorus and sulfur atoms follow the octet rule. However, there are times where phosphorus follows the ten-electrons rule, and sulfur follows either the ten-electrons or the twelve-electrons rule. Why do they happen and what types of quantum orbitals do they involve? Take a look again at all the molecules that contain sulfur or phosphorus in this lab, explain how some of your initial shapes might be wrong. Correct any wrong assumptions.

(Internet search phrase: “formal charges of (*name of ion or molecule*)” or go to <http://www.chem.plu.edu/pchem/chem342/FormChg.htm> as the starting point of your research.)

Conclusion:

Summarize what you have learned from this lab.

*Note:

2. For SiH₄, use a black ball to represent Si.
3. For PH₃, use a blue ball to represent P.
6. For H₂S, use a red ball to represent S.
8. For O₃, use a black ball to represent the central atom.
13. For SO₂, use a black ball to represent S.
14. For PF₅, use a brown ball (or any ball with 5 holes) to represent P.
15. For SCl₆, use a yellow ball (or any ball with 6 holes) to represent S.
16. For PO₄³⁻, use a black ball to represent P if it follows the octet rule. If P follows 10e⁻ rule, use the brown ball (or any ball with 5 holes) to represent P.
17. For SO₄²⁻, use a black ball to represent S if it follows the octet rule. If S follows 10e⁻ rule, use the brown ball (or any ball with 5 holes) to represent S. If S follows 12e⁻ rule, use the yellow ball (or any ball with 6 holes) to represent S.
18. For SCl₄, use a black ball to represent S if it follows the octet rule. If S follows 10e⁻ rule, use the brown ball (or any ball with 5 holes) to represent S. If S follows 12e⁻ rule, use the yellow ball (or any ball with 6 holes) to represent S.
19. For IF₂¹⁻, use a black ball to represent I if it follows the octet rule. If I follows 10e⁻ rule, use the brown ball (or any ball with 5 holes) to represent I. If I follows 12e⁻ rule, use the grey ball (or any ball with 6 holes) to represent I.
20. For BrF₅, use the brown ball (or any ball with 5 holes) to represent Br if it follows 10e⁻ rule. If Br follows 12e⁻ rule, use the grey ball (or any ball with 6 holes) to represent Br.
21. For XeCl₄, use a black ball to represent Xe if it follows the octet rule. If Xe follows 10e⁻ rule, use the brown ball (or any ball with 5 holes) to represent Xe. If Xe follows 12e⁻ rule, use the grey ball (or any ball with 6 holes) to represent Xe.
22. For BrCl₃, use a black ball to represent Br if it follows the octet rule. If Br follows 10e⁻ rule, use the brown ball (or any ball with 5 holes) to represent Br. If Br follows 12e⁻ rule, use the grey ball (or any ball with 6 holes) to represent Br.