

## Lab #8: Chemical Equilibrium

### Objective:

To observe and explain the effect of an applied stress on chemical systems at equilibrium by applying the Le Châtelier's Principle.

### Background Information:

In this lab, the effect of applying stresses to a variety of chemical systems at equilibrium will be explored. The equilibrium systems to be studied are given below:

- Saturated Sodium Chloride Solution:**  $\text{NaCl}_{(s)} \rightleftharpoons \text{Na}^+_{(aq)} + \text{Cl}^-_{(aq)}$
- Acidified Chromate Solution:**  $2 \text{CrO}_4^{2-}_{(aq)} + 2 \text{H}^+_{(aq)} \rightleftharpoons \text{Cr}_2\text{O}_7^{2-}_{(aq)} + \text{H}_2\text{O}_{(l)}$   
(yellow) (acid) (red)
- Ammonia Solution (with phenolphthalein):**  $\text{NH}_3_{(aq)} + \text{H}_2\text{O}_{(l)} \rightleftharpoons \text{NH}_4^+ + \text{OH}^-_{(aq)} + \text{Heat}$   
(NH<sub>3</sub> clear in phenolphthalein) (OH<sup>-</sup> is pink in phenolphthalein)
- Cobalt (II) Chloride Solution:**  $\text{Co}(\text{H}_2\text{O})_6^{2+}_{(aq)} + 4 \text{Cl}^-_{(aq)} + \text{Heat} \rightleftharpoons \text{CoCl}_4^{2-}_{(aq)} + 6 \text{H}_2\text{O}_{(l)}$   
(reddish-pink) (blue)
- Iron (III) Thiocyanate Solution:**  $\text{Fe}^{3+}_{(aq)} + \text{SCN}^-_{(aq)} \rightleftharpoons \text{Fe}(\text{SCN})^-_{(aq)} + \text{Heat}$   
(pale yellow) (colourless) (deep red)

By observing the changes that occur (colour changes, precipitate formation, etc.) the direction of a particular shift may be determined. Such shifts may then be explained by carefully examining the effect of the applied stress as dictated by Le Châtelier's Principle.

### Hypotheses:

Predict the *shift in equilibrium* **and** *possible observations* for the following situations.

- Saturated Sodium Chloride Solution:**  $\text{NaCl}_{(s)} \rightleftharpoons \text{Na}^+_{(aq)} + \text{Cl}^-_{(aq)}$ 
  - Adding HCl<sub>(aq)</sub>
- Acidified Chromate Solution:**  $2 \text{CrO}_4^{2-}_{(aq)} + 2 \text{H}^+_{(aq)} \rightleftharpoons \text{Cr}_2\text{O}_7^{2-}_{(aq)} + \text{H}_2\text{O}_{(l)}$   
(yellow) (acid) (red)
  - Adding HNO<sub>3</sub>(aq) first
  - Adding NaOH<sub>(aq)</sub> after in the same test tube (H<sup>+</sup> + OH<sup>-</sup> forms H<sub>2</sub>O)
- Ammonia Solution (with phenolphthalein):**  $\text{NH}_3_{(aq)} + \text{H}_2\text{O}_{(l)} \rightleftharpoons \text{NH}_4^+ + \text{OH}^-_{(aq)} + \text{Heat}$   
(NH<sub>3</sub> clear in phenolphthalein) (OH<sup>-</sup> is pink in phenolphthalein)
  - Adding NH<sub>4</sub>Cl<sub>(s)</sub>
  - Decreasing Temperature
  - Increasing Temperature
- Cobalt (II) Chloride Solution:**  $\text{Co}(\text{H}_2\text{O})_6^{2+}_{(aq)} + 4 \text{Cl}^-_{(aq)} + \text{Heat} \rightleftharpoons \text{CoCl}_4^{2-}_{(aq)} + 6 \text{H}_2\text{O}_{(l)}$   
(reddish-pink) (blue)
  - Adding H<sub>2</sub>O<sub>(l)</sub>
  - Adding HCl<sub>(aq)</sub>
  - Adding AgNO<sub>3</sub>(aq)
  - Decreasing Temperature
  - Increasing Temperature
- Iron (III) Thiocyanate Solution:**  $\text{Fe}^{3+}_{(aq)} + \text{SCN}^-_{(aq)} \rightleftharpoons \text{Fe}(\text{SCN})^-_{(aq)} + \text{Heat}$   
(pale yellow) (colourless) (deep red)
  - Adding Fe(NO<sub>3</sub>)<sub>3</sub>(aq)
  - Adding KSCN<sub>(aq)</sub>
  - Adding K<sub>2</sub>HPO<sub>4</sub>(aq) (HPO<sub>4</sub><sup>2-</sup>(aq) and Fe<sup>3+</sup>(aq) forms a complex FeHPO<sub>4</sub><sup>+</sup>(aq))
  - Decreasing Temperature

**Materials:**

14 Small Test Tubes	Stirring Rods	6 M HNO <sub>3</sub> (aq)	0.1 M AgNO <sub>3</sub> (aq)
Test Tube Racks	Medicine Droppers	6 M NaOH (aq)	0.1 M Fe(NO <sub>3</sub> ) <sub>3</sub> (aq) <sup>‡</sup>
Bunsen Burner	Ice and Water	Diluted NH <sub>3</sub> (aq) <sup>*</sup>	0.1 M KSCN (aq) <sup>‡</sup>
2 Beakers (250 mL)	Saturated NaCl (aq)	NH <sub>4</sub> Cl (s)	0.1 M K <sub>2</sub> HPO <sub>4</sub> (aq)
Test Tube Clamps	Phenolphthalein	0.1 M K <sub>2</sub> CrO <sub>4</sub> (aq) <sup>*</sup>	0.1 M K <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub> (aq) <sup>*</sup>
Scoopula	12 M HCl (aq)	0.1 M CoCl <sub>2</sub> (aq) <sup>†</sup>	0.2 M KCl (aq) <sup>†</sup>

<sup>\*</sup> To make the acidify chromate solutions, combine equal volumes of 0.1 M of K<sub>2</sub>CrO<sub>4</sub> (aq) and 0.1 M K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> (aq). Add about one-tenth of the original volume of 6 M HNO<sub>3</sub> (aq). The initial colour of the final solution should be orange.

<sup>\*</sup> Prepare a stock solution by adding 4 drops of concentrated 6 M NH<sub>3</sub> (aq) and 3 drops of phenolphthalein to a medium beaker, then topping it up with 250 mL of water. Mix with a stirring rod. The entire class can share this stock solution.

<sup>†</sup> CoCl<sub>2</sub> (s) forms complex ions, Co(H<sub>2</sub>O)<sub>6</sub><sup>2+</sup> (aq) and CoCl<sub>4</sub><sup>2-</sup> (aq), when dissolves in water. An equilibrium cobalt (II) chloride solution with its complex ions can be prepared by mixing twice equal volume of 0.1 M CoCl<sub>2</sub> (aq) with 0.2 M KCl (aq).

<sup>\*</sup> Prepare a stock solution of iron (III) thiocyanate by adding 10 mL of 0.1 M FeCl<sub>3</sub> (aq) and 10 mL of 0.1 M KSCN (aq) to a medium beaker, then topping it up with 100 mL of water. Mix with a stirring rod. The entire class can share this stock solution.

**Procedure:**

Record all observations on your report form. These should include, but not be limited to, colour changes and precipitates. Note that solution volumes are approximate for all reactions below. Dispose of all chemical waste in the plastic container in the fume hood.

**Safety**

All of the acids and bases used in this experiment (HCl (aq), HNO<sub>3</sub> (aq) and NaOH (aq)) can cause chemical burns. In particular, concentrated 12 M HCl and 6 M HNO<sub>3</sub> are extremely dangerous! If any of these chemicals spill on you, immediately rinse the affected area under running water and notify your instructor. Also note that direct contact with silver nitrate (AgNO<sub>3</sub>) will cause dark discolourations to appear on your skin. These spots will eventually fade after repeated rinses in water. Finally, in Parts 3, 4 and 5, you will be heating a solution in a test tube directly in a Bunsen Burner flame. If the solution is overheated it will splatter out of the tube, so be careful not to point the tube towards anyone while heating.

**Part 1: Saturated Sodium Chloride Solution (Demo: 1 test tube)**

- Place 2 mL (about the width of a thumb nail) of saturated NaCl (aq) into a small test tube. Record the initial observation.
- (Demo)** Observe as your instructor *carefully* add the concentrated 12 M HCl (aq) drop-wise to the test tube containing the saturated NaCl (aq) until a distinct change occurs. Record your observations.

**Part 2: Acidified Chromate Solution (1 test tube)**

- Place about 2 mL (about the width of a thumb nail) of 0.1 M K<sub>2</sub>CrO<sub>4</sub> (aq) / K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> (aq) into a small test tube. Label this as test tube #1. Record the initial observation.
- Carefully* add concentrated 6 M HNO<sub>3</sub> (aq) drop-wise to this solution. Record your observations.
- To the same test tube, *carefully* add concentrated 6 M NaOH (aq) drop-wise until a distinct change occurs. Record your observations.

Part 3: Aqueous Ammonia Solution (2 test tubes)

1. To each of the two test tubes, place 2 mL (about the width of a thumb nail) of the previously diluted  $\text{NH}_3$  (aq) solution with phenolphthalein. Label these test tubes #2 & #3. Record the initial observation.
2. Add a medium scoop of  $\text{NH}_4\text{Cl}$  (s) powder to test tube #2 until it is completely dissolve. Record your observations.
3. Take test tube #3 and add a few drops of phenolphthalein and place it in the beaker full of ice. Wait until a distinct change occurs. Be patient. Record your observations.
4. Take the previous test tube #3 from the beaker of ice. Heat the test tube over the Bunsen Burner flame by firmly holding it with the test tube clamp, and waft it back and forth through the flame (to prevent overheating and “bumping”) for about 30 seconds, or, until a distinct change occurs. Record your observations. Place the test tube in a beaker of room temperature water when finished. Remove the test tube clamp from the test tube only when it is cooled.

Part 4: Cobalt (II) Chloride Solution (4 test tubes)

1. To each of the four test tubes, place 2 mL (about the width of a thumb nail) of 0.1 M  $\text{CoCl}_2$  (aq). Label these test tubes 4 to 7. Record the initial observation.
2. The solution in test tube #4 remains untouched. It is a control for comparison with other tubes.
3. Add about 4 mL of water to one of the test tube #5. Record your observations.
4. **(Demo)** Observe as your instructor *carefully* add the concentrated 12 M  $\text{HCl}$  (aq) drop-wise to a test tube containing the 0.1 M  $\text{CoCl}_2$  (aq). Record your observations.
5. Add 0.1 M  $\text{AgNO}_3$  (aq) drop-wise to the fourth test tube #6 until any distinct changes occur. Record your observations.
6. Take the last test tube, #7 and place it in the beaker full of ice. Wait until a distinct change occurs. Be patient. Record your observations.
7. Take the previous test tube #7 from the beaker of ice. Heat the test tube over the Bunsen Burner flame by firmly holding it with the test tube clamp, and waft it back and forth through the flame (to prevent overheating and “bumping”) for about 30 seconds, or, until a distinct change occurs. Record your observations. Place the test tube in a beaker of room temperature water when finished. Remove the test tube clamp from the test tube only when it is cooled.

Part 5: Iron (III) Thiocyanate Solution (5 test tubes)

1. To each of the five test tubes, place 2 mL (about the width of a thumb nail) of the iron (III) thiocyanate solution into 4 small test tubes. Label these test tubes 8 to 12. Record the initial observation.
2. The solution in test tube #8 remains untouched. It is a control for comparison with other tubes.
3. Add 2 mL of 0.1 M of  $\text{Fe}(\text{NO}_3)_3$  (aq) to test tube #9. Record your observations.
4. Add 2 mL of 0.1 M of  $\text{KSCN}$  (aq) to one of the test tube #10. Record your observations.
5. Add 2 mL of 0.1 M of  $\text{K}_2\text{HPO}_4$  (aq) to one of the test tube #11. Record your observations.
6. Take the last test tube, #12 and place it in the beaker full of ice. Wait until a distinct change occurs. Be patient. Record your observations.

**Observations:****Part 1: Saturated Sodium Chloride Solution (Demo)**

Original Saturated Sodium Chloride Solution	
<b>Demo:</b> Addition of Concentrated HCl <sub>(aq)</sub>	

**Part 2: Acidified Chromate Solution**

Original Acidified Chromate Solution	
Test Tube #1: Addition of Concentrated HNO <sub>3</sub> (aq)	
Test Tube #1: Further Addition of Concentrated NaOH (aq)	

**Part 3: Aqueous Ammonia Solution**

Original Aqueous Ammonia Solution	
Test Tube #2: Addition of NH <sub>4</sub> Cl (s)	
Test Tube #3: Placed in Ice	
Test Tube #3: Heating with Bunsen Burner	

**Part 4: Cobalt (II) Chloride Solution**

Test Tube #4: Original CoCl <sub>2</sub> Solution	
Test Tube #5: Addition of H <sub>2</sub> O (l)	
<b>Demo:</b> Addition of Concentrated HCl (aq)	
Test Tube #6: Addition of AgNO <sub>3</sub> (aq)	
Test Tube #7: Placed in Ice	
Test Tube #7: Heating with Bunsen Burner	

**Part 5: Iron (III) Thiocyanate Solution**

Test Tube #8: Original Iron (III) Thiocyanate Solution	
Test Tube #9: Addition of Fe(NO <sub>3</sub> ) <sub>3</sub> (aq)	
Test Tube #10: Addition of KSCN (aq)	
Test Tube #11: Addition of K <sub>2</sub> HPO <sub>4</sub> (aq)	
Test Tube #12: Placed in Ice	

**Analysis**

In a detailed manner, explain how the stresses affect each of the five equilibrium systems. Be sure to use the Le Châtelier's Principle and any other chemical concepts to account for all your observations.

**Evaluation:**

How did your hypotheses compare to the observations and the analysis? Account for any discrepancies.

**Conclusion:**

Summarize what you have learned from this lab.